

AD-A097 009

ROCKWELL INTERNATIONAL ANAHEIM CA AUTONETICS STRATEG--ETC F/6 7/4  
CATHODIC ELECTROCHROMISM OF LUTETIUM DIPHthalOCYANINE FILMS.(U)

MAR 81 M M NICHOLSON, F A PIZZARELLO

N00014-77-C-0636

UNCLASSIFIED

C60-785/201

NL

1 OF 1  
AD-A  
097009



END  
DATE  
FILMED  
4-8  
DTIC

# LEVEL II

(12)  
C80-785/201

OFFICE OF NAVAL RESEARCH

Contract N00014-77-C-0636

Task No. NR 359-667

TECHNICAL REPORT NO. 3

Cathodic Electrochromism of Lutetium Diphthalocyanine Films

by

M. M. Nicholson and F. A. Pizzarello

Prepared for Publication

in the

Journal of the Electrochemical Society

Autonetics Strategic Systems Division  
Rockwell International  
Anaheim, California

March 1981

Reproduction in whole or in part is permitted for any  
purpose of the United States Government.

This document has been approved for public release  
and sale; its distribution is unlimited.

DTIC  
ELECTE  
MAR 30 1981  
S D F

81 3 30 008

AD A 097009

DTIC FILE COPY

UNCLASSIFIED

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER	2. GOVT ACCESSION NO.	3. RECIPIENT'S CATALOG NUMBER
	AD-A097009	
4. TITLE (and Subtitle)		5. TYPE OF REPORT & PERIOD COVERED
Cathodic Electrochromism of Lutetium Diphthalocyanine Films.		Technical 1 Sept 1980 - 31 Aug 1980
7. AUTHOR(s)		6. PERFORMING OR REPORT NUMBER
M. M. Nicholson and F. A. Pizzarello		C80-785/201 TR-3
9. PERFORMING ORGANIZATION NAME AND ADDRESS		8. CONTRACT OR GRANT NUMBER (if any)
Autonetics Strategic Systems Division- Rockwell International 3370 Miraloma Avenue, Anaheim, CA 92803		NO0014-77-C-0636
11. CONTROLLING OFFICE NAME AND ADDRESS		10. PROGRAM ELEMENT PROJECT, TASK AREA & WORK UNIT NUMBERS
Office of Naval Research 800 N. Quincy Arlington, VA 22217		
14. MONITORING AGENCY NAME & ADDRESS (if different from Controlling Office)		12. REPORT DATE
12 23		March 1981
		13. NUMBER OF PAGES
		17
		15. SECURITY CLASS. (of this report)
		Unclassified
		15a. DECLASSIFICATION DOWNGRADING SCHEDULE
16. DISTRIBUTION STATEMENT (of this Report)		
This document has been approved for public release and sale; its distribution is unlimited.		
17. DISTRIBUTION STATEMENT (of the abstract entered in Block 20, if different from Report)		
18. SUPPLEMENTARY NOTES		
Prepared for publication in the Journal of the Electrochemical Society.		
19. KEY WORDS (Continue on reverse side if necessary and identify by block number)		
diphthalocyanines electrochromism cathodic reduction cations		
20. ABSTRACT (Continue on reverse side if necessary and identify by block number)		
Cathodic electrochromism of lutetium diphthalocyanine films on insulating substrates was investigated by a moving-boundary technique. Reduction occurred by injection of electrons from a gold contact and cations from a liquid electrolyte. Light blue products formed with aqueous or organic solutions of alkali metal salts, while a dark violet product formed with aqueous hydrochloric acid. The dark material had a charge-carrier mobility of $8 \times 10^{-7}$ cm <sup>2</sup> /V-sec and an estimated bulk resistivity of 1,800 ohm-cm. Reduced forms of the dye appear to be solid cation conductors.		

DD FORM 1473

1 JAN 73

EDITION OF NOV 65 IS OBSOLETE

UNCLASSIFIED

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

391827

## CATHODIC ELECTROCHROMISM OF LUTETIUM DIPHthalOCYANINE FILMS

M. M. Nicholson\* and F. A. Pizzarello

Rockwell International, Autonetics Strategic Systems Division,  
Anaheim, California 92803

## ABSTRACT

Cathodic electrochromism of lutetium diphthalocyanine films on insulating substrates was investigated by a moving-boundary technique. Reduction occurred by injection of electrons from a gold contact and cations from a liquid electrolyte. Light blue products formed with aqueous or organic solutions of alkali metal salts, while a dark violet product formed with aqueous hydrochloric acid. The dark material had a charge-carrier mobility of  $2 \times 10^{-7}$  cm<sup>2</sup>/V-sec and an estimated bulk resistivity of 1,800 ohm-cm. Reduced forms of the dye appear to be solid cation conductors.

Accession For  
 THIS CASE  
 FILED IN  
 CHARGE OF  
 THE OFFICE OF  
 THE ATTORNEY GENERAL

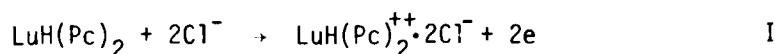
RECEIVED  
 DEPT. OF JUSTICE  
 DIVISION OF INVESTIGATION  
 APR 10 1968

A

\*Electrochemical Society Active Member.

**Key words:** diphthalocyanines, electrochromism, cathodic reduction, cations.

Previous papers from this laboratory have reported investigations of anodic processes in lutetium diphthalocyanine films (1-4). Generally, these are 2-electron oxidations in which the color of the dye changes from green to red. The electrochromic reaction was shown to occur with migration of anions from the aqueous electrolyte into the organic solid. With a chloride solution, it is represented\* by



A solid-state moving-boundary technique was used to determine the faradaic  $n$  values, ionic mobilities, and approximate bulk resistivities in the red phases formed with chloride and sulfate ions (1). The anion mechanism was confirmed in a subsequent tracer study (2), and its dependences on ambient water vapor and oxygen were reported recently (4).

Lutetium diphthalocyanine also undergoes cathodic reactions. As a result, it is a multicolor electrochromic material of considerable interest for display applications (5-7). Colors of the reduced films can range from blue-green through blue, dark blue, and violet. Reference 5 gives absorption spectra recorded at 0.1-V intervals on a tin-oxide supported film under essentially open-circuit conditions in both anodic and cathodic ranges. Spectral characteristics as well as Munsell and CIE color coordinates are summarized in References 5 and 6.

---

\*  $\text{H}_2\text{Pc}$  is the usual abbreviation for metal-free phthalocyanine;  $\text{Pc}^-$  is the anion  $\text{C}_{32}\text{H}_{16}\text{N}_8^-$ , and  $\text{LuH(Pc)}_2$  is commonly called lutetium diphthalocyanine.

This paper describes the results of cathodic moving-boundary experiments on lutetium diphthalocyanine films supported by insulating substrates. The electrolytes were aqueous HCl, aqueous KCl, KI in acetonitrile (AN), and LiCl in dimethyl sulfoxide (DMSO). Again, the solid reaction products appeared to be ionic conductors. The propagated color was light blue with the metal salt solutions and violet with HCl.

### Experimental

Lutetium diphthalocyanine films were prepared by vacuum sublimation of the dye onto sapphire or Mylar strips 1.25 cm wide and approximately 5 cm long. The weight of dye per  $\text{cm}^2$  was determined from the optical density (OD) of the green film at the 670-nm absorption maximum. The geometric film thickness was estimated from this weight and the approximate crystal density. It ranged from 530 to 1,060 Å. Details of the dye synthesis and specimen preparation were given previously (1). The other materials were reagent-grade chemicals. The organic solvents were dried by allowing them to stand over a molecular sieve.

A sputtered gold contact was applied at the upper end of the dye film, and the lower end was immersed in the liquid electrolyte. The counter electrode was silver foil. A drawing of the moving-boundary cell is given in Reference 1. The cell was assembled in a helium-atmosphere glove box, where the oxygen level did not exceed a few ppm. Constant cathodic currents of 0.2 to 2  $\mu\text{A}$  were applied with a PAR 173 galvanostat, and the total applied voltage was monitored with a Keithley 610C electrometer and a strip-chart recorder. The boundary-propagation distance was measured by viewing the film against a back-lighted screen of millimeter graph paper under  $\sim 2\times$  magnification. The

electric field in the color-converted film could be determined by raising the electrolyte level in successive increments and observing the corresponding voltage decreases. In several cases, the specimens were examined after propagation by energy-dispersive x-ray spectroscopy (EDS), using an ETEC Autoscan scanning electron microscope. Comparisons of lutetium, potassium, and chlorine levels in different regions of the films were made in this way.

### Results and Discussion

Visual Observations.-- The color change always began at the dye/electrolyte interface and travelled upward toward the electronic contact. This behavior is consistent with an ion-injection mechanism. It differs from that of Yamana's qualitative experiments, wherein the blue coloration of erbium diphthalocyanine films on glass began at the metal contact (8). In our experience, color-converted diphthalocyanine films are prone to crack and peel away from glass. If this occurred in Yamana's cell, the dye could have become wet with electrolyte by capillary action. Cathodic reaction then would have occurred in the upper part of the film.

The adhesion problem was mitigated, though not fully avoided, in the present work by using a sapphire or Mylar substrate.\* With aqueous electrolytes, the cathodic boundary propagation tended not to proceed as uniformly as it did for the anodic green-to-red process (1), but the boundary velocity and field were measurable in a number of experiments, as shown by Fig. 1-3. Occasionally, the boundary was too irregular for quantitative use.

---

\* The dye film adheres well to tin oxide, but that substrate cannot be used in moving-boundary studies because it is electronically conductive.

An additional difficulty arose with the organic electrolytes. Although the initial green form of the dye was insoluble in several organic liquids, including AN and DMSO, the blue reduced form was soluble. This led to poor electrical contact during the boundary propagation and prevented determination of the field by changing the electrolyte level.

The propagated color depended on the electrolyte cation. With HCl, the product was dark violet. With KCl, KI, and LiCl, only a light blue appeared. In contrast, we have observed that anodic boundary propagation with different anions yields a single red color (4).

Electrical Parameters.-- The results for aqueous and nonaqueous solutions are summarized in Table I. Although the total applied currents were in the microampere range, the estimated current densities through the cross sections of the films were 15 to 200 mA/cm<sup>2</sup>. This range includes current densities that would be used in potentiostatic switching of a dipthalocyanine electrochromic display (7).

The apparent number  $n$  of electrons transferred per molecule of dye was calculated from the current, the area converted per unit time, and the optical density of the green film. Plots for evaluating the boundary velocity are shown in Fig. 1 and 2. With aqueous KCl,  $n$  was  $1.9 \pm 0.2$ . With HCl, the apparent  $n$  was considerably larger, and on Specimen 1, it was shown to increase systematically with the applied current. The large apparent  $n$ 's in the acid solution probably were caused by a side reaction forming elemental hydrogen. The apparent  $n$  at 0.4  $\mu$ A was 5.2 for Specimen 1 and 12 for Specimen 2. One could surmise from this comparison that the sapphire substrate was more conducive than Mylar to proton transport beneath the dye film, but further investigation would be needed to verify such a mechanism. Extrapolation of the apparent  $n$  to zero current yielded a value of 3.5 for

the Mylar-supported film contacted by 1.2 M HCl. Fewer data were recorded for nonaqueous electrolytes because of the blue-solubility problem, but an  $n$  value of 1.1 was obtained with both Mylar- and sapphire-supported films contacted by LiCl-DMSO.

A carrier mobility  $\mu$  in the converted film was found from the relationship

$$\mu = v/\mathcal{E} \quad (1)$$

where  $v$  is the boundary velocity and  $\mathcal{E}$  is the electric field determined by raising the electrolyte level (1). Figure 3 shows plots for the evaluation of  $\mathcal{E}$  in the Mylar-supported film propagated from aqueous HCl. The average mobility was  $8 \times 10^{-7} \text{ cm}^2/\text{V-sec}$ . Since the current and the approximate film thickness were known, the bulk resistivity  $\rho_b$  of the blue phase could be estimated from Eq. [2], where  $I$  is the current density.

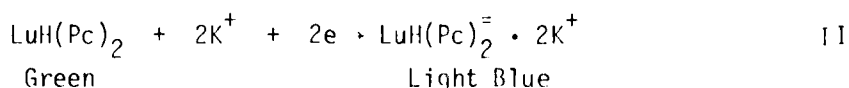
$$\rho_b = \mathcal{E}/I \quad (2)$$

The average resistivity for Specimen 1 was 1,800 ohm-cm. This is similar to the resistivities of 1,300 and 2,000 ohm-cm for the red oxidation products containing chloride and sulfate ions (1).

Interpretations.-- The violet phase propagated from an aqueous HCl interface appears to be an ionic conductor. This is suggested by the magnitude of the carrier mobility, which is compared in Table II with those of related solid

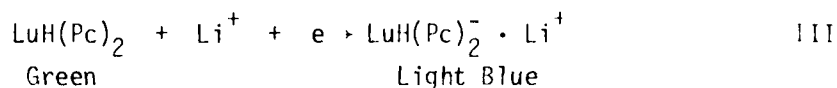
electrolytes. Dependence of the reduction-product color on the electrolyte cation also is consistent with a cation-injection mechanism. Moreover, EDS analysis confirmed the presence of potassium, and the absence of chlorine, in blue areas propagated from KCl, while neither of these elements was found in green areas of the same plates.

The reaction for aqueous KCl may be written



This process is analogous to the anodic oxidation utilizing two chloride ions and forming the divalent organic cation.

The measurements with LiCl in DMSO indicated a 1-electron reduction, which might be represented by

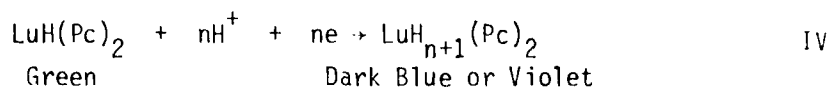


With dissolved lutetium diphthalocyanine, Corker, Grant, and Clecak also reported cathodic formation of a radical anion (11). The electrolyte was tetrabutylammonium fluoborate in dimethylformamide, and the working electrode was platinum. Reduction beyond the 1-electron stage with LiCl in the moving-boundary cell might have been prevented by separation of the two phthalocyanine rings to form LiHPc and LuPc. Such a process could account for the higher solubility in organic solvents.

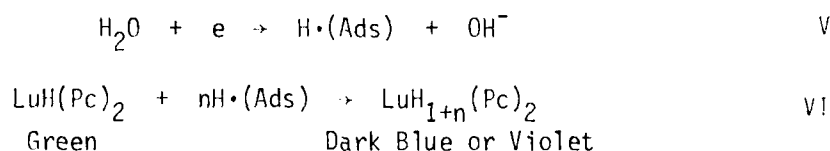
Further work is needed to characterize the reduction products. It is significant, however, that dark blue or violet colors never were obtained with metal cations in the moving-boundary experiments, even at a current density as high as  $200 \text{ mA/cm}^2$ . An acidic electrolyte is required to form the dark product in insulator-supported films. Yet dark blue and violet are readily obtained

with aqueous KCl when the dye is on semiconductive tin oxide (5,6). We propose the following mechanisms to account for this difference in the electrochromic behavior.

With an adequate supply of protons, reduction of the dye can proceed by a double-injection mechanism similar to that of Reaction II:



All or part of the n+1 hydrogen atoms may be ionized in the solid state. When the supply of protons is insufficient, the dye may be converted at a tin oxide electrode by a coupled chemical reaction rather than direct electron transfer:



The reactive intermediate is represented as an adsorbed hydrogen atom because a species with such characteristics is known to form on tin oxide at cathodic potentials (12,13). Furthermore, a pH increase accompanies cathodic formation of a blue product on tin oxide in aqueous KCl, and hydrogen ions are released to the solution on its anodic oxidation (14).

The dark blue and violet forms cannot yet be distinguished on the basis of the n values. One of our previous investigations provides evidence, from correlation of absorption spectra with approximate open-circuit potentials of the reduced films, that the dark blue and violet colors may involve two discrete oxidation states of the dye (5). We now estimate that n is 2 or 3 for the dark blue state, and 3 or 4 for the violet.

Moskalev and Kirin originally proposed field-induced ionization to  $\text{H}^+$  and  $\text{Lu(Pc)}_2^-$  to account for the cathodic electrochromism of lutetium diphthalocyanine

films (15). The present study has shown that these color changes are due, instead, to faradaic processes. Although the acidic hydrogen may not be firmly bound to the organic structure in rare-earth diphthalocyanines (16), it is the electrolyte cation that plays a vital role in the electrochromism at cathodic potentials.

#### Acknowledgment

This work was supported in part by the Office of Naval Research.

# REFERENCES

1. M. M. Nicholson and F. A. Pizzarello, This Journal, 126, 1490 (1979).
2. F. A. Pizzarello and M. M. Nicholson, J. Electron. Mater., 9, 231 (1980).
3. M. M. Nicholson and F. A. Pizzarello, This Journal, 127, 821 (1980).
4. F. A. Pizzarello and M. M. Nicholson, *ibid.*, 127, 2617 (1980).
5. M. M. Nicholson and R. V. Galiardi, Final Report, Contract N62269-76-C-0574, AD-A039596, May 1977.
6. M. M. Nicholson and R. V. Galiardi, SID International Symposium Digest, IX, 24 (1978).
7. M. M. Nicholson, F. A. Pizzarello, and T. J. La Chapelle, Final Report, Contract N00014-79-C-0434, ONR-CR399-005-1F, June 1980.
8. M. Yamana, Oyo Butsuri, 48, 441 (1979).
9. B. W. Faughnan, R. S. Crandall, and M. A. Lampert, Appl. Phys. Lett., 27, 275 (1975).
10. P. E. Childs and T. K. Halstead, Mater. Res. Bull., 13, 609 (1978).
11. G. A. Corker, B. Grant, and M. J. Clecak, This Journal, 126, 1339 (1979).
12. H. A. Laitinen, C. A. Vincent, and T. M. Bednarski, *ibid.*, 115, 1024 (1968).
13. M. M. Nicholson and R. V. Galiardi, Unpublished work.
14. M. M. Nicholson, F. A. Pizzarello, R. V. Galiardi, and G. A. Layman, Final Report, Contracts F49620-77-C-0074 and F49620-77-C-0104, June 1980.
15. P. N. Moskalev and I. S. Kirin, Russ. J. Phys. Chem., 46, 1019 (1972).
16. K. Kasuga, M. Tsutsui, R. C. Petterson, K. Tatsumi, N. Van Opdenbosch, G. Pepe, and E. F. Meyer, Jr., J. Am. Chem. Soc., 102, 4835 (1980).

Table I. Results of Cathodic Boundary Propagation in Lutetium Diphthalocyanine Films

Specimen	Initial OD at 670 nm	Substrate	Solvent	Electrolyte	Current (.A)	Color Formed	Apparent n	(cm <sup>2</sup> /V-sec) x 10 <sup>7</sup>
1	1.15	Mylar	Water	1.2 M HCl	0 0.2 0.4 0.8 1.4	Violet	3.5* 4.7 5.2 7.6 8.8	-- 13 4.1 6.2 8.4
2	1.02	Sapphire	Water	1 M HCl	0.4	Violet	12	--
3	0.82	Sapphire	Water	1 M KCl	0.5	Light blue	1.7	--
4	0.85	Sapphire	Water	1 M KCl	1	Light blue	2.1	--
5	0.85	Sapphire	Water	1 M KCl	2	Light blue	1.8	--
6	0.69	Sapphire	AN	0.1 M KI	0.5	Light blue	~3	--
7	0.58	Mylar	DMSO	~2M LiCl	0.5	Light blue	1.1	--
8	1.02	Sapphire	DMSO	~2M LiCl	0.5	Light blue	1.1	--

\*By extrapolation of n vs i to i = 0.

Table II. Solid-State Ionic Mobilities Near 25

Material	Form	Mobile Ion	Mobility (cm <sup>2</sup> /V-sec) x 10 <sup>6</sup>	Reference
Anodic oxidation product of LuH(Pc) <sub>2</sub>	Thin film	Cl <sup>-</sup> = SO <sub>4</sub>	4 4	1 1
Cathodic reduction product of LuH(Pc) <sub>2</sub>	Thin film	H <sup>+</sup>	0.8	This investigation
H <sub>x</sub> WO <sub>3</sub>	Thin film	H <sup>+</sup>	0.8	9
HWO <sub>2</sub> PO <sub>4</sub> · 4H <sub>2</sub> O	Polycrystalline	H <sup>+</sup>	8	10

### Illustrations

1. Dependence of boundary propagation distance on time with aqueous 1.2 M HCl electrolyte. Specimen 1.
2. Dependence of boundary propagation distance on time with aqueous 1 M KCl electrolyte. Specimens 3, 4, and 5.
3. Plots for determination of electric field in the blue film with aqueous 1.2 M HCl electrolyte. Specimen 1.

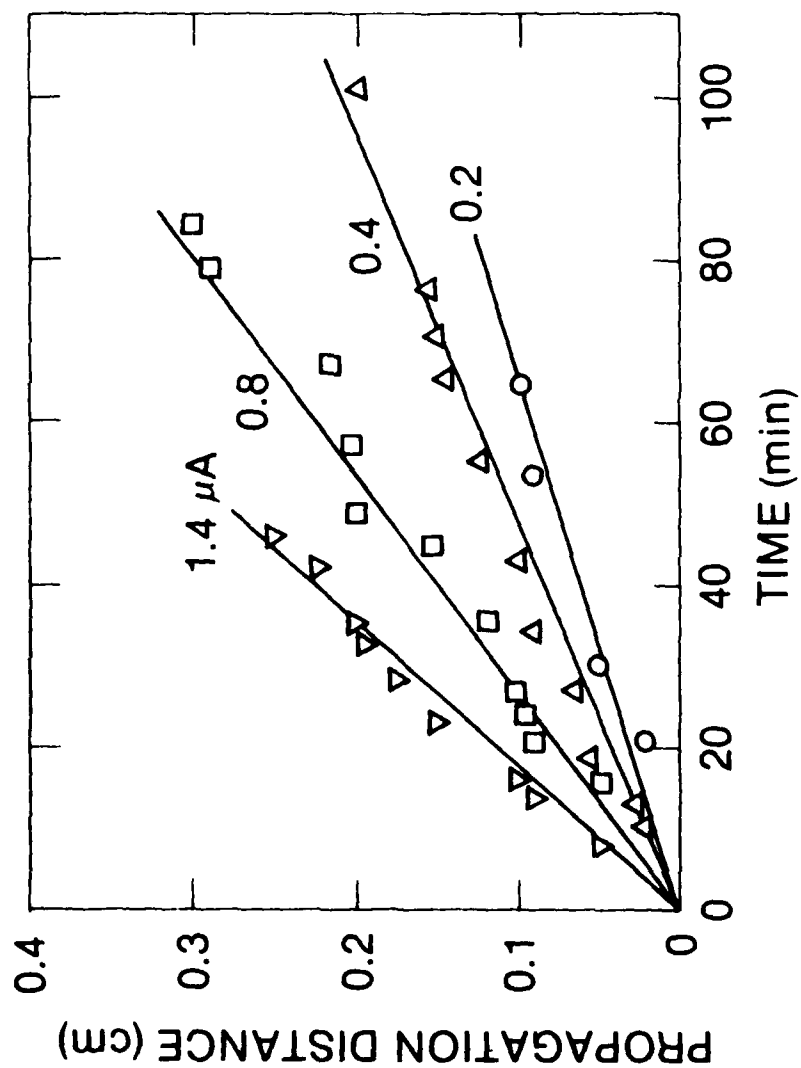


FIG. 1

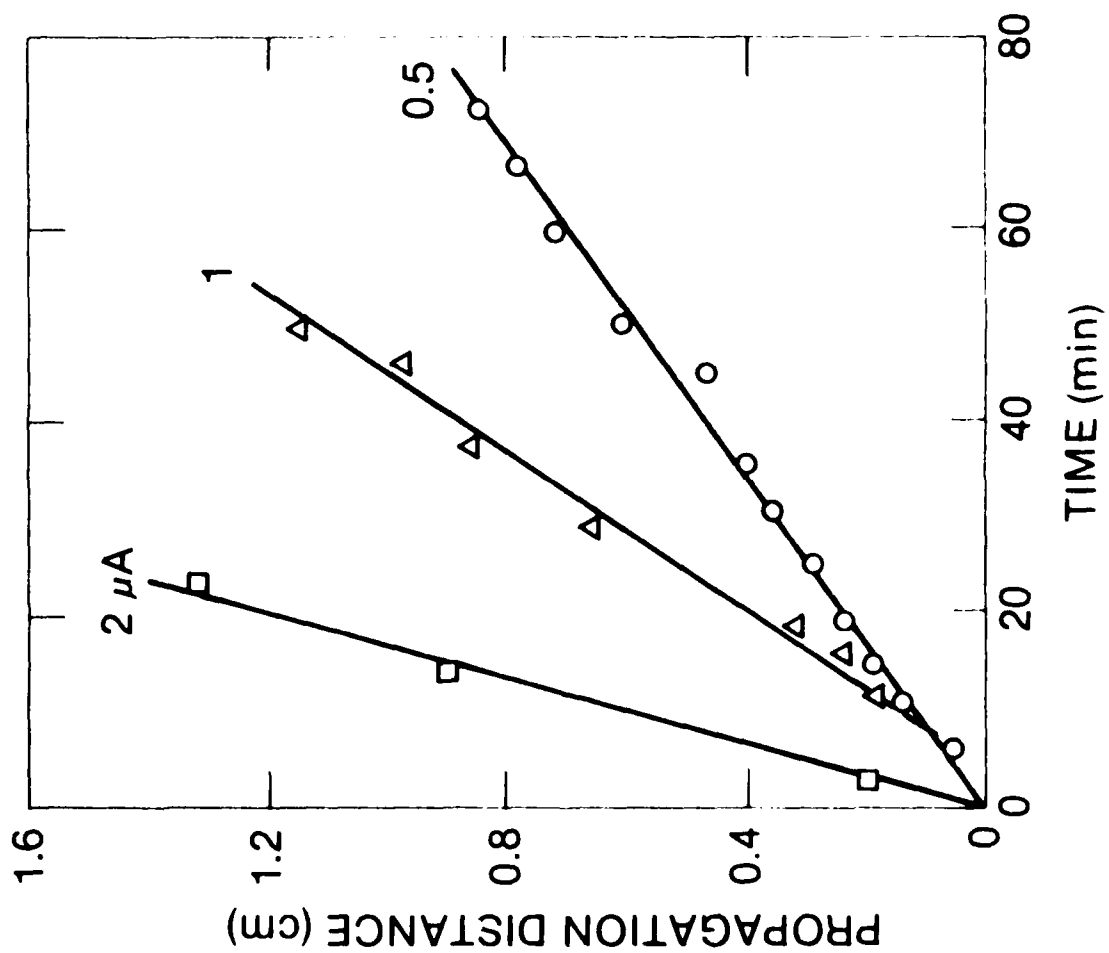


FIG. 2

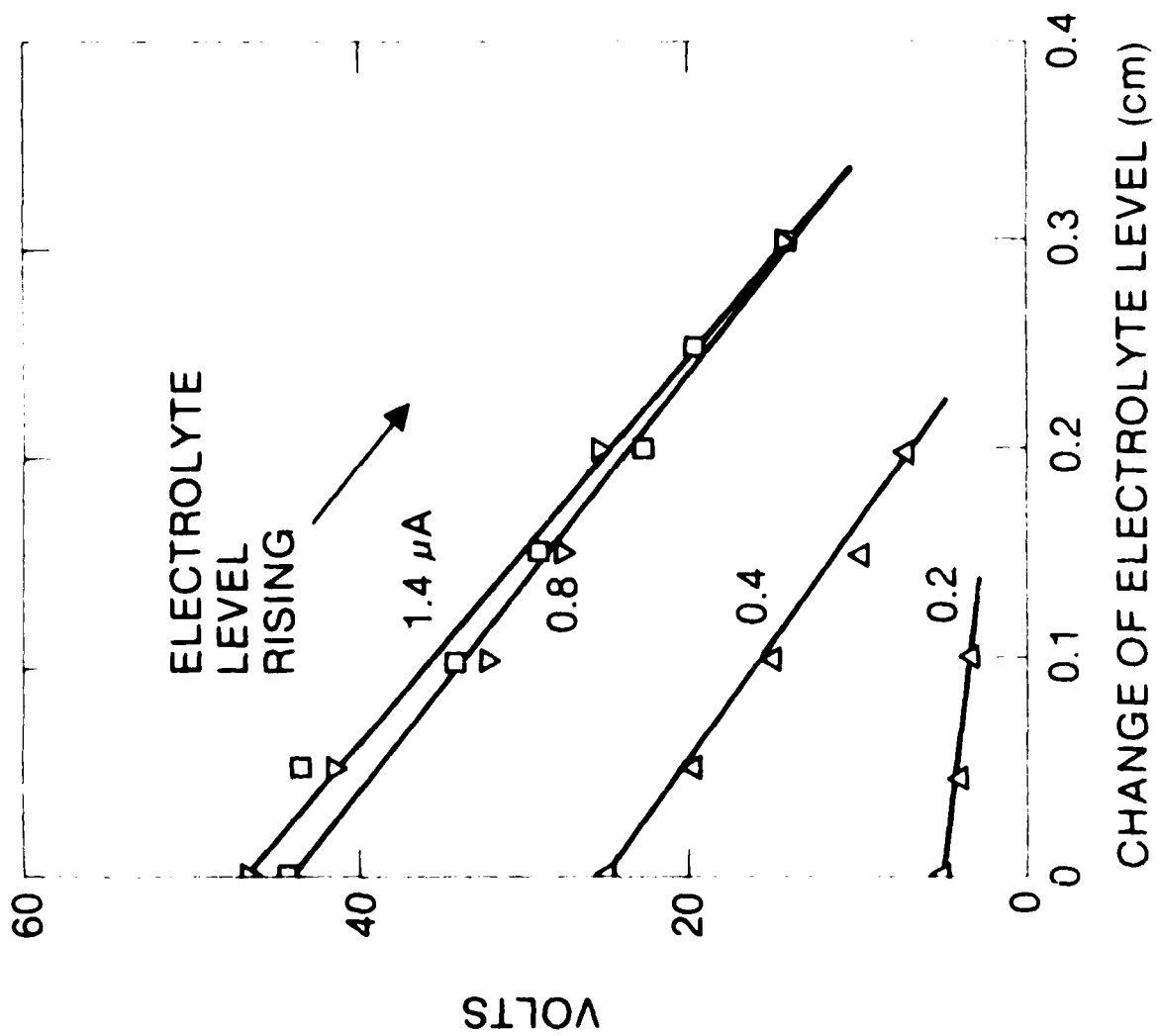


FIG. 3

TECHNICAL REPORT DISTRIBUTION LIST, GEN

	<u>No.</u> <u>Copies</u>		<u>No.</u> <u>Copies</u>
Office of Naval Research Attn: Code 472 800 North Quincy Street Arlington, Virginia 22217	2	U.S. Army Research Office Attn: CRD-AA-IP P.O. Box 1211 Research Triangle Park, N.C. 27709	1
ONR Branch Office Attn: Dr. George Sandoz 536 S. Clark Street Chicago, Illinois 60605	1	Naval Ocean Systems Center Attn: Mr. Joe McCartney San Diego, California 92152	1
ONR Area Office Attn: Scientific Dept. 715 Broadway New York, New York 10003	1	Naval Weapons Center Attn: Dr. A. B. Amster, Chemistry Division China Lake, California 93555	1
ONR Western Regional Office 1030 East Green Street Pasadena, California 91106	1	Naval Civil Engineering Laboratory Attn: Dr. R. W. Drisko Port Hueneme, California 93401	1
ONR Eastern/Central Regional Office Attn: Dr. L. H. Peebles Building 114, Section D 666 Summer Street Boston, Massachusetts 02210	1	Department of Physics & Chemistry Naval Postgraduate School Monterey, California 93940	1
Director, Naval Research Laboratory Attn: Code 6100 Washington, D.C. 20390	1	Dr. A. L. Slafkosky Scientific Advisor Commandant of the Marine Corps (Code RD-1) Washington, D.C. 20380	1
The Assistant Secretary of the Navy (RE&S) Department of the Navy Room 4E736, Pentagon Washington, D.C. 20350	1	Office of Naval Research Attn: Dr. Richard S. Miller 800 N. Quincy Street Arlington, Virginia 22217	1
Commander, Naval Air Systems Command Attn: Code 310C (H. Rosenwasser) Department of the Navy Washington, D.C. 20360	1	Naval Ship Research and Development Center Attn: Dr. G. Bosmajian, Applied Chemistry Division Annapolis, Maryland 21401	1
Defense Technical Information Center Building 5, Cameron Station Alexandria, Virginia 22314	12	Naval Ocean Systems Center Attn: Dr. S. Yamamoto, Marine Sciences Division San Diego, California 91232	1
Dr. Fred Saalfeld Chemistry Division, Code 6100 Naval Research Laboratory Washington, D.C. 20375	1	Mr. John Boyle Materials Branch Naval Ship Engineering Center Philadelphia, Pennsylvania 19112	1

TECHNICAL REPORT DISTRIBUTION LIST, GENNo.  
Copies

Dr. Rudolph J. Marcus  
Office of Naval Research  
Scientific Liaison Group  
American Embassy  
APO San Francisco 96503

1

Mr. James Kelley  
DTNSRDC Code 2803  
Annapolis, Maryland 21402

1

TECHNICAL REPORT DISTRIBUTION LIST, 359

	<u>No. Copies</u>		<u>No. Copies</u>
Dr. Paul Delahay Department of Chemistry New York University New York, New York 10003	1	Dr. P. J. Hendra Department of Chemistry University of Southampton Southampton SO9 5NH United Kingdom	1
Dr. E. Yeager Department of Chemistry Case Western Reserve University Cleveland, Ohio 44106	1	Dr. Sam Perone Department of Chemistry Purdue University West Lafayette, Indiana 47907	1
Dr. D. N. Bennion Department of Chemical Engineering Brigham Young University Provo, Utah 84602	1	Dr. Royce W. Murray Department of Chemistry University of North Carolina Chapel Hill, North Carolina 27514	1
Dr. R. A. Marcus Department of chemistry California Institute of Technology Pasadena, California 91125	1	Naval Ocean Systems Center Attn: Technical Library San Diego, California 92152	1
Dr. J. J. Auborn Bell Laboratories Murray Hill, New Jersey 07974	1	Dr. C. E. Mueller The Electrochemistry Branch Materials Division, Research & Technology Department Naval Surface Weapons Center White Oak Laboratory Silver Spring, Maryland 20910	1
Dr. Adam Heller Bell Laboratories Murray Hill, New Jersey 07974	1	Dr. G. Goodman Globe-Union Incorporated 5757 North Green Bay Avenue Milwaukee, Wisconsin 53201	1
Dr. T. Eatan Lockheed Missiles & Space Co., Inc. P.O. Box 504 Sunnyvale, California 94088	1	Dr. J. Boechler Electrochimica Corporation Attention: Technical Library 2485 Charleston Road Mountain View, California 94040	1
Dr. Joseph Singer, Code 302-1 NASA-Lewis 21000 Brookpark Road Cleveland, Ohio 44135	1	Dr. P. E. Schmidt Department of Chemistry Oakland University Rochester, Michigan 48063	1
Dr. B. Brummer EIC Incorporated 55 Chapel Street Newton, Massachusetts 02458	1	Dr. H. Richtel Chemistry Department Rensselaer Polytechnic Institute Troy, New York 12181	1
Library P. R. Mallory and Company, Inc. Northwest Industrial Park Burlington, Massachusetts 01803	1		

TECHNICAL REPORT DISTRIBUTION LIST, 359

	<u>No.</u> <u>Copies</u>		<u>No.</u> <u>Copies</u>
Dr. A. B. Ellis Chemistry Department University of Wisconsin Madison, Wisconsin 53706	1	Dr. R. P. Van Duyne Department of Chemistry Northwestern University Evanston, Illinois 60201	1
Dr. M. Wrighton Chemistry Department Massachusetts Institute of Technology Cambridge, Massachusetts 02139	1	Dr. B. Stanley Pons Department of Chemistry University of Alberta Edmonton, Alberta CANADA T6G 2G2	1
Larry E. Flew Naval Weapons Support Center Code 30736, Building 2906 Crane, Indiana 47522	1	Dr. Michael J. Weaver Department of Chemistry Michigan State University East Lansing, Michigan 48824	1
S. Rube DOE (STOE) 600 F Street Washington, D.C. 20545	1	Dr. R. David Rauh EIC Corporation 55 Chapel Street Newton, Massachusetts 02158	1
Dr. Aaron Wold Brown University Department of Chemistry Providence, Rhode Island 02192	1	Dr. J. David Margerum Research Laboratories Division Hughes Aircraft Company 3011 Malibu Canyon Road Malibu, California 90265	1
Dr. R. C. Chudacek McGraw-Edison Company Edison Battery Division Post Office Box 28 Bloomfield, New Jersey 07003	1	Dr. Martin Fleischmann Department of Chemistry University of Southampton Southampton SO9 5NH England	1
Dr. A. J. Bard University of Texas Department of Chemistry Austin, Texas 78712	1	Dr. Janet Ostervoung Department of Chemistry State University of New York at Buffalo Buffalo, New York 14214	1
<del>Dr. M. M. Nicholson Electronics Research Center Rockwell International 3370 Arrialoma Avenue Anaheim, California</del>	<del>1</del>	Dr. R. A. Ostervoung Department of Chemistry State University of New York at Buffalo Buffalo, New York 14214	<del>1</del>
Dr. Donald W. Ernst Naval Surface Weapons Center Code R-33 White Oak Laboratory Silver Spring, Maryland 20910	1	Mr. James R. Moden Naval Underwater Systems Center Code 3632 Newport, Rhode Island 02840	1

TECHNICAL REPORT DISTRIBUTION LIST, 359

<u>No.</u> <u>Copies</u>	<u>No.</u> <u>Copies</u>
Dr. R. Nowak Naval Research Laboratory Code 6130 Washington, D.C. 20375 1	Dr. John Kincaid Department of the Navy Strategic Systems Project Office Room 901 Washington, DC 20376 1
Dr. John F. Houlihan Shenango Valley Campus Pennsylvania State University Sharon, Pennsylvania 16146 1	M. L. Robertson Manager, Electrochemical Power Sonics Division Naval Weapons Support Center Crane, Indiana 47522 1
Dr. M. G. Sceats Department of Chemistry University of Rochester Rochester, New York 14627 1	Dr. Elton Cairns Energy & Environment Division Lawrence Berkeley Laboratory University of California Berkeley, California 94720 1
Dr. D. F. Shriver Department of Chemistry Northwestern University Evanston, Illinois 60201 1	Dr. Bernard Spielvogel U.S. Army Research Office P.O. Box 12211 Research Triangle Park, NC 27709 1
Dr. D. H. Whitmore Department of Materials Science Northwestern University Evanston, Illinois 60201 1	Dr. Denton Elliott Air Force Office of Scientific Research Bldg. 104 Bolling AFB Washington, DC 20332 1
Dr. Alan Perlick Department of Chemistry The University Southampton, SO9 5NH England 1	
Dr. A. Himy NAVSEA-5433 NC #4 2541 Jefferson Davis Highway Arlington, Virginia 20362 1	

END

DATE  
FILMED

4-8-1

DTIC